

## 3.4 Ions and Ionic Compounds

### LEARNING OBJECTIVES

1. Know how ions form.
2. Learn the characteristic charges that ions have.
3. Construct a proper formula for an ionic compound.
4. Generate a proper name for an ionic compound.

So far, we have discussed elements and compounds that are electrically neutral. They have the same number of electrons as protons, so the negative charges of the electrons is balanced by the positive charges of the protons. However, this is not always the case. Electrons can move from one atom to another; when they do, species with overall electric charges are formed. Such species are called **ions**. Species with overall positive charges are termed **cations**, while species with overall negative charges are called **anions**. Remember that ions are formed only when *electrons* move from one atom to another; a proton never moves from one atom to another. Compounds formed from positive and negative ions are called ionic compounds.

Individual atoms can gain or lose electrons. When they do, they become *monatomic ions*. When atoms gain or lose electrons, they usually gain or lose a characteristic number of electrons and so take on a characteristic overall charge. [Table 3.6 "Monatomic Ions of Various Charges"](#) lists some common ions in terms of how many electrons they lose (making cations) or gain (making anions). There are several things to notice about the ions in [Table 3.6 "Monatomic Ions of Various Charges"](#). First, each element that forms cations is a metal, except for one (hydrogen), while each element that forms anions is a nonmetal. This is actually one of the chemical properties of metals and nonmetals: metals tend to form cations, while nonmetals tend to form anions. Second, most atoms form ions of a single characteristic charge. When sodium atoms form ions, they always form a 1+ charge, never a 2+ or 3+ or even 1– charge. Thus, if you commit the information in [Table 3.6 "Monatomic Ions of Various Charges"](#) to memory, you will always know what charges most atoms form. (In [Chapter 9 "Chemical Bonds"](#), we will discuss *why* atoms form the charges they do.)

Table 3.6 Monatomic Ions of Various Charges

	H <sup>+</sup>
	Na <sup>+</sup>
	K <sup>+</sup>
	Rb <sup>+</sup>
	Ag <sup>+</sup>
Ions formed by losing a single electron	Au <sup>+</sup>
Ions formed by losing two electrons	Mg <sup>2+</sup>

	$\text{Ca}^{2+}$
	$\text{Sr}^{2+}$
	$\text{Fe}^{2+}$
	$\text{Co}^{2+}$
	$\text{Ni}^{2+}$
	$\text{Cu}^{2+}$
	$\text{Zn}^{2+}$
	$\text{Sn}^{2+}$
	$\text{Hg}^{2+}$
	$\text{Pb}^{2+}$
Ions formed by losing three electrons	$\text{Sc}^{3+}$
	$\text{Fe}^{3+}$
	$\text{Co}^{3+}$
	$\text{Ni}^{3+}$
	$\text{Au}^{3+}$
	$\text{Al}^{3+}$
	$\text{Cr}^{3+}$
Ions formed by losing four electrons	$\text{Ti}^{4+}$
	$\text{Sn}^{4+}$



	$\text{Pb}^{4+}$
Ions formed by gaining a single electron	$\text{F}^{-}$
	$\text{Cl}^{-}$
	$\text{Br}^{-}$
	$\text{I}^{-}$
Ions formed by gaining two electrons	$\text{O}^{2-}$
	$\text{S}^{2-}$
	$\text{Se}^{2-}$
Ions formed by gaining three electrons	$\text{N}^{3-}$
	$\text{P}^{3-}$

Third, there are some exceptions to the previous point. A few elements, all metals, can form more than one possible charge. For example, iron atoms can form 2+ cations or 3+ cations. Cobalt is another element that can form more than one possible charged ion (2+ and 3+), while lead can form 2+ or 4+ cations. Unfortunately, there is little understanding which two charges a metal atom may take, so it is best to just memorize the possible charges a particular element can have.

Note the convention for indicating an ion. The magnitude of the charge is listed as a right superscript next to the symbol of the element. If the charge is a single positive or negative one, the number 1 is not written; if the magnitude of the charge is greater than 1, then the number is written *before* the + or – sign. An element symbol without a charge written next to it is assumed to be the uncharged atom.

Naming an ion is straightforward. For a cation, simply use the name of the element and add the word *ion* (or if you want to be more specific, add *cation*) after the element's name. So  $\text{Na}^+$  is the sodium ion;  $\text{Ca}^{2+}$  is the calcium ion. If the element has more than one possible charge, the value of the charge comes after the element name and before the word *ion*. Thus,  $\text{Fe}^{2+}$  is the iron two ion, while  $\text{Fe}^{3+}$  is the iron three ion. In print, we use roman numerals in parentheses to represent the charge on the ion, so these two iron ions would be represented as the iron(II) cation and the iron(III) cation, respectively.

For a monatomic anion, use the stem of the element name and append the suffix *-ide* to it, and then add *ion*. This is similar to how we named molecular compounds. Thus,  $\text{Cl}^-$  is the chloride ion, and  $\text{N}^{3-}$  is the nitride ion.

### EXAMPLE 6

Name each species.

1.  $\text{O}^{2-}$
2.  $\text{Co}$
3.  $\text{Co}^{2+}$

Solution

1. This species has a 2- charge on it, so it is an anion. Anions are named using the stem of the element name with the suffix *-ide* added. This is the oxide anion.
2. Because this species has no charge, it is an atom in its elemental form. This is cobalt.



3. In this case, there is a 2+ charge on the atom, so it is a cation. We note from [Table 3.6 "Monatomic Ions of Various Charges"](#) that cobalt cations can have two possible charges, so the name of the ion must specify which charge the ion has. This is the cobalt(II) cation.

### *Test Yourself*

Name each species.

1.  $\text{P}^{3-}$
2.  $\text{Sr}^{2+}$

### *Answers*

1. the phosphide anion
2. the strontium cation

Chemical formulas for ionic compounds are called **ionic formulas**. A proper ionic formula has a cation and an anion in it; an ionic compound is never formed between two cations only or two anions only. The key to writing proper ionic formulas is simple: the total positive charge must balance the total negative charge. Because the charges on the ions are characteristic, sometimes we have to have more than one of a cation or an anion to balance the overall positive and negative charges. It is conventional to use the lowest ratio of ions that are needed to balance the charges.

For example, consider the ionic compound between  $\text{Na}^+$  and  $\text{Cl}^-$ . Each ion has a single charge, one positive and one negative, so we need only

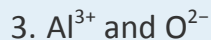
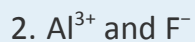
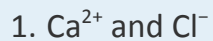
one ion of each to balance the overall charge. When writing the ionic formula, we follow two additional conventions: (1) write the formula for the cation first and the formula for the anion next, but (2) do not write the charges on the ions. Thus, for the compound between  $\text{Na}^+$  and  $\text{Cl}^-$ , we have the ionic formula  $\text{NaCl}$  ([Figure 3.5 "NaCl = Table Salt"](#)). The formula  $\text{Na}_2\text{Cl}_2$  also has balanced charges, but the convention is to use the lowest ratio of ions, which would be one of each. (Remember from our conventions for writing formulas that we don't write a 1 subscript if there is only one atom of a particular element present.) For the ionic compound between magnesium cations ( $\text{Mg}^{2+}$ ) and oxide anions ( $\text{O}^{2-}$ ), again we need only one of each ion to balance the charges. By convention, the formula is  $\text{MgO}$ .

For the ionic compound between  $\text{Mg}^{2+}$  ions and  $\text{Cl}^-$  ions, we now consider the fact that the charges have different magnitudes, 2+ on the magnesium ion and 1- on the chloride ion. To balance the charges with the lowest number of ions possible, we need to have two chloride ions to balance the charge on the one magnesium ion. Rather than write the formula  $\text{MgClCl}$ , we combine the two chloride ions and write it with a 2 subscript:  $\text{MgCl}_2$ .

What is the formula  $\text{MgCl}_2$  telling us? There are two chloride ions in the formula. Although chlorine as an element is a diatomic molecule,  $\text{Cl}_2$ , elemental chlorine is not part of this ionic compound. The chlorine is in the form of a negatively charged *ion*, not the neutral *element*. The 2 subscript is in the ionic formula because we need two  $\text{Cl}^-$  ions to balance the charge on one  $\text{Mg}^{2+}$  ion.

### EXAMPLE 7

Write the proper ionic formula for each of the two given ions.



### Solution

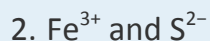
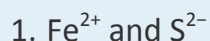
1. We need two  $\text{Cl}^-$  ions to balance the charge on one  $\text{Ca}^{2+}$  ion, so the proper ionic formula is  $\text{CaCl}_2$ .

2. We need three  $\text{F}^-$  ions to balance the charge on the  $\text{Al}^{3+}$  ion, so the proper ionic formula is  $\text{AlF}_3$ .

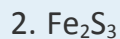
3. With  $\text{Al}^{3+}$  and  $\text{O}^{2-}$ , note that neither charge is a perfect multiple of the other. This means we have to go to a least common multiple, which in this case will be six. To get a total of 6+, we need two  $\text{Al}^{3+}$  ions; to get 6-, we need three  $\text{O}^{2-}$  ions. Hence the proper ionic formula is  $\text{Al}_2\text{O}_3$ .

### Test Yourself

Write the proper ionic formulas for each of the two given ions.



### Answers





Naming ionic compounds is simple: combine the name of the cation and the name of the anion, in both cases omitting the word *ion*. *Do not use numerical prefixes if there is more than one ion necessary to balance the charges*. NaCl is sodium chloride, a combination of the name of the cation (sodium) and the anion (chloride). MgO is magnesium oxide. MgCl<sub>2</sub> is magnesium chloride—*not* magnesium dichloride.

In naming ionic compounds whose cations can have more than one possible charge, we must also include the charge, in parentheses and in roman numerals, as part of the name. Hence FeS is iron(II) sulfide, while Fe<sub>2</sub>S<sub>3</sub> is iron(III) sulfide. Again, no numerical prefixes appear in the name. The number of ions in the formula is dictated by the need to balance the positive and negative charges.

### EXAMPLE 8

Name each ionic compound.

1. CaCl<sub>2</sub>
2. AlF<sub>3</sub>
3. Co<sub>2</sub>O<sub>3</sub>

Solution

1. Using the names of the ions, this ionic compound is named calcium chloride. *It is not calcium(II) chloride* because calcium forms only one cation when it forms an ion, and it has a characteristic charge of 2+.
2. The name of this ionic compound is aluminum fluoride.

3. We know that cobalt can have more than one possible charge; we just need to determine what it is. Oxide always has a 2<sup>-</sup> charge, so with three oxide ions, we have a total negative charge of 6<sup>-</sup>. This means that the two cobalt ions have to contribute 6<sup>+</sup>, which for two cobalt ions means that each one is 3<sup>+</sup>. Therefore, the proper name for this ionic compound is cobalt(III) oxide.

### *Test Yourself*

Name each ionic compound.

1.  $\text{Sc}_2\text{O}_3$
2.  $\text{AgCl}$

### *Answers*

1. scandium oxide
2. silver chloride

How do you know whether a formula—and by extension, a name—is for a molecular compound or for an ionic compound? Molecular compounds form between nonmetals and nonmetals, while ionic compounds form between metals and nonmetals. The periodic table ([Figure 3.2 "A Simple Periodic Table"](#)) can be used to determine which elements are metals and nonmetals.

There also exists a group of ions that contain more than one atom. These are called **polyatomic ions**. [Table 3.7 "Common Polyatomic Ions"](#) lists the formulas, charges, and names of some common polyatomic ions.

Only one of them, the ammonium ion, is a cation; the rest are anions.

Most of them also contain oxygen atoms, so sometimes they are referred to as *oxyanions*. Some of them, such as nitrate and nitrite, and sulfate and sulfite, have very similar formulas and names, so care must be taken to get the formulas and names correct. Note that the -ite polyatomic ion has one less oxygen atom in its formula than the -ate ion but with the same ionic charge.

Table 3.7 Common Polyatomic Ions

Name	Formula and Charge	Name	Formula and Charge
ammonium	$\text{NH}_4^+$	hydroxide	$\text{OH}^-$
acetate	$\text{C}_2\text{H}_3\text{O}_2^-$ , or $\text{CH}_3\text{COO}^-$	nitrate	$\text{NO}_3^-$
bicarbonate (hydrogen carbonate)	$\text{HCO}_3^-$	nitrite	$\text{NO}_2^-$
bisulfate (hydrogen sulfate)	$\text{HSO}_4^-$	peroxide	$\text{O}_2^{2-}$
carbonate	$\text{CO}_3^{2-}$	perchlorate	$\text{ClO}_4^-$
chlorate	$\text{ClO}_3^-$	phosphate	$\text{PO}_4^{3-}$
chromate	$\text{CrO}_4^{2-}$	sulfate	$\text{SO}_4^{2-}$
cyanide	$\text{CN}^-$	sulfite	$\text{SO}_3^{2-}$
dichromate	$\text{Cr}_2\text{O}_7^{2-}$	triiodide	$\text{I}_3^-$

The naming of ionic compounds that contain polyatomic ions follows the same rules as the naming for other ionic compounds: simply combine the name of the cation and the name of the anion. Do not use

numerical prefixes in the name if there is more than one polyatomic ion; the only exception to this is if the name of the ion itself contains a numerical prefix, such as dichromate or triiodide.

Writing the formulas of ionic compounds has one important difference. If more than one polyatomic ion is needed to balance the overall charge in the formula, enclose the formula of the polyatomic ion in parentheses and write the proper numerical subscript to the right and *outside* the parentheses. Thus, the formula between calcium ions,  $\text{Ca}^{2+}$ , and nitrate ions,  $\text{NO}_3^-$ , is properly written  $\text{Ca}(\text{NO}_3)_2$ , not  $\text{CaNO}_{32}$  or  $\text{CaN}_2\text{O}_6$ . Use parentheses where required. The name of this ionic compound is simply calcium nitrate.

### EXAMPLE 9

Write the proper formula and give the proper name for each ionic compound formed between the two listed ions.

1.  $\text{NH}_4^+$  and  $\text{S}^{2-}$
2.  $\text{Al}^{3+}$  and  $\text{PO}_4^{3-}$
3.  $\text{Fe}^{2+}$  and  $\text{PO}_4^{3-}$

#### Solution

1. Because the ammonium ion has a 1+ charge and the sulfide ion has a 2- charge, we need two ammonium ions to balance the charge on a single sulfide ion. Enclosing the formula for the ammonium ion in parentheses, we have  $(\text{NH}_4)_2\text{S}$ . The compound's name is ammonium sulfide.
2. Because the ions have the same magnitude of charge, we need only one of each to balance the charges. The formula is  $\text{AlPO}_4$ , and the name of the compound is aluminum phosphate.

3. Neither charge is an exact multiple of the other, so we have to go to the least common multiple of 6. To get 6+, we need three iron(II) ions, and to get 6−, we need two phosphate ions. The proper formula is  $\text{Fe}_3(\text{PO}_4)_2$ , and the compound's name is iron(II) phosphate.

### *Test Yourself*

Write the proper formula and give the proper name for each ionic compound formed between the two listed ions.

1.  $\text{NH}_4^+$  and  $\text{PO}_4^{3-}$
2.  $\text{Co}^{3+}$  and  $\text{NO}_2^-$

### *Answers*

1.  $(\text{NH}_4)_3\text{PO}_4$ , ammonium phosphate
2.  $\text{Co}(\text{NO}_2)_3$ , cobalt(III) nitrite

## **Food and Drink App: Sodium in Your Food**

The element sodium, at least in its ionic form as  $\text{Na}^+$ , is a necessary nutrient for humans to live. In fact, the human body is approximately 0.15% sodium, with the average person having one-twentieth to one-tenth of a kilogram in their body at any given time, mostly in fluids outside cells and in other bodily fluids.

Sodium is also present in our diet. The common table salt we use on our foods is an ionic sodium compound. Many processed foods also contain significant amounts of sodium added to them as a variety of ionic compounds. Why are sodium compounds used so much? Usually sodium compounds are inexpensive, but, more importantly, most ionic



sodium compounds dissolve easily. This allows processed food manufacturers to add sodium-containing substances to food mixtures and know that the compound will dissolve and distribute evenly throughout the food. Simple ionic compounds such as sodium nitrite ( $\text{NaNO}_2$ ) are added to cured meats, such as bacon and deli-style meats, while a compound called sodium benzoate is added to many packaged foods as a preservative. [Table 3.8 "Some Sodium Compounds Added to Food"](#) is a partial list of some sodium additives used in food. Some of them you may recognize after reading this chapter. Others you may not recognize, but they are all ionic sodium compounds with some negatively charged ion also present.

Table 3.8 Some Sodium Compounds Added to Food

Sodium Compound	Use in Food
Sodium acetate	preservative, acidity regulator
Sodium adipate	food acid
Sodium alginate	thickener, vegetable gum, stabilizer, gelling agent, emulsifier
Sodium aluminum phosphate	acidity regulator, emulsifier
Sodium aluminosilicate	anticaking agent
Sodium ascorbate	antioxidant
Sodium benzoate	preservative
Sodium bicarbonate	mineral salt
Sodium bisulfite	preservative, antioxidant

<b>Sodium Compound</b>	<b>Use in Food</b>
Sodium carbonate	mineral salt
Sodium carboxymethylcellulose	emulsifier
Sodium citrates	food acid
Sodium dehydroacetate	preservative
Sodium erythorbate	antioxidant
Sodium erythorbin	antioxidant
Sodium ethyl para-hydroxybenzoate	preservative
Sodium ferrocyanide	anticaking agent
Sodium formate	preservative
Sodium fumarate	food acid
Sodium gluconate	stabilizer
Sodium hydrogen acetate	preservative, acidity regulator
Sodium hydroxide	mineral salt
Sodium lactate	food acid
Sodium malate	food acid
Sodium metabisulfite	preservative, antioxidant, bleaching agent
Sodium methyl para-hydroxybenzoate	preservative



Sodium Compound	Use in Food
Sodium nitrate	preservative, color fixative
Sodium nitrite	preservative, color fixative
Sodium orthophenyl phenol	preservative
Sodium propionate	preservative
Sodium propyl para-hydroxybenzoate	preservative
Sodium sorbate	preservative
Sodium stearyl lactylate	emulsifier
Sodium succinates	acidity regulator, flavor enhancer
Sodium salts of fatty acids	emulsifier, stabilizer, anticaking agent
Sodium sulfite	mineral salt, preservative, antioxidant
Sodium sulfite	preservative, antioxidant
Sodium tartrate	food acid
Sodium tetraborate	preservative

The use of so many sodium compounds in prepared and processed foods has alarmed some physicians and nutritionists. They argue that the average person consumes too much sodium from his or her diet. The average person needs only about 500 mg of sodium every day; most people consume more than this—up to 10 times as much. Some studies have implicated increased sodium intake with high blood pressure; newer studies suggest that the link is questionable. However, there has



been a push to reduce the amount of sodium most people ingest every day: avoid processed and manufactured foods, read labels on packaged foods (which include an indication of the sodium content), don't oversalt foods, and use other herbs and spices besides salt in cooking.

Nutrition Facts			
Serving Size 8 oz (227 g/8 oz)			
Servings Per Container About 3			
Amount Per Serving			
Calories	180	Calories from Fat 60	
		% Daily Value*	
Total Fat	6g		10%
Saturated Fat	1g		5%
Trans Fat	0g		
Cholesterol	5mg		2%
Sodium	75mg		3%
Total Carbohydrate	26g		9%
Dietary Fiber	5g		19%
Sugars	11g		
Protein	8g		
Vitamin A 60% • Vitamin C 70%			
Calcium 8% • Iron 10%			
* Percent Daily Values are based on a 2,000 calorie diet. Your daily values may be higher or lower depending on your calorie needs.			
	Calories	2,000	2,500
Total Fat	Less than	65g	80g
Sat Fat	Less than	20g	25g
Cholesterol	Less than	300mg	300mg
Sodium	Less than	2,400mg	2,400mg
Total Carbohydrate		300g	375g
Dietary Fiber		25g	30g
Calories per gram:			
Fat 9 • Carbohydrate 4 • Protein 4			

*Food labels include the amount of sodium per serving. This particular label shows that there are 75 mg of sodium in one serving of this particular food item.*

## KEY TAKEAWAYS

- Ions form when atoms lose or gain electrons.
- Ionic compounds have positive ions and negative ions.
- Ionic formulas balance the total positive and negative charges.
- Ionic compounds have a simple system of naming.
- Groups of atoms can have an overall charge and make ionic compounds.

## EXERCISES

1. Explain how cations form.
2. Explain how anions form.
3. Give the charge each atom takes when it forms an ion. If more than one charge is possible, list both.
  - a. K
  - b. O
  - c. Co
4. Give the charge each atom takes when it forms an ion. If more than one charge is possible, list both.
  - a. Ca
  - b. I
  - c. Fe
5. Give the charge each atom takes when it forms an ion. If more than one charge is possible, list both.

- a. Ag
  - b. Au
  - c. Br
6. Give the charge each atom takes when it forms an ion. If more than one charge is possible, list both.
- a. S
  - b. Na
  - c. H
7. Name the ions from Exercise 3.
8. Name the ions from Exercise 4.
9. Name the ions from Exercise 5.
10. Name the ions from Exercise 6.
11. Give the formula and name for each ionic compound formed between the two listed ions.
- a.  $\text{Mg}^{2+}$  and  $\text{Cl}^-$
  - b.  $\text{Fe}^{2+}$  and  $\text{O}^{2-}$
  - c.  $\text{Fe}^{3+}$  and  $\text{O}^{2-}$
12. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{K}^+$  and  $\text{S}^{2-}$
- b.  $\text{Ag}^+$  and  $\text{Br}^-$
- c.  $\text{Sr}^{2+}$  and  $\text{N}^{3-}$

13. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Cu}^{2+}$  and  $\text{F}^-$
- b.  $\text{Ca}^{2+}$  and  $\text{O}^{2-}$
- c.  $\text{K}^+$  and  $\text{P}^{3-}$

14. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Na}^+$  and  $\text{N}^{3-}$
- b.  $\text{Co}^{2+}$  and  $\text{I}^-$
- c.  $\text{Au}^{3+}$  and  $\text{S}^{2-}$

15. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{K}^+$  and  $\text{SO}_4^{2-}$
- b.  $\text{NH}_4^+$  and  $\text{S}^{2-}$
- c.  $\text{NH}_4^+$  and  $\text{PO}_4^{3-}$

16. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Ca}^{2+}$  and  $\text{NO}_3^-$

- b.  $\text{Ca}^{2+}$  and  $\text{NO}_2^-$
- c.  $\text{Sc}^{3+}$  and  $\text{C}_2\text{H}_3\text{O}_2^-$

17. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Pb}^{4+}$  and  $\text{SO}_4^{2-}$
- b.  $\text{Na}^+$  and  $\text{I}_3^-$
- c.  $\text{Li}^+$  and  $\text{Cr}_2\text{O}_7^{2-}$

18. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{NH}_4^+$  and  $\text{N}^{3-}$
- b.  $\text{Mg}^{2+}$  and  $\text{CO}_3^{2-}$
- c.  $\text{Al}^{3+}$  and  $\text{OH}^-$

19. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Ag}^+$  and  $\text{SO}_3^{2-}$
- b.  $\text{Na}^+$  and  $\text{HCO}_3^-$
- c.  $\text{Fe}^{3+}$  and  $\text{ClO}_3^-$

20. Give the formula and name for each ionic compound formed between the two listed ions.

- a.  $\text{Rb}^+$  and  $\text{O}_2^{2-}$
- b.  $\text{Au}^{3+}$  and  $\text{HSO}_4^-$



c.  $\text{Sr}^{2+}$  and  $\text{NO}_2^-$

21. What is the difference between  $\text{SO}_3$  and  $\text{SO}_3^{2-}$ ?

22. What is the difference between  $\text{NO}_2$  and  $\text{NO}_2^-$ ?

## ANSWERS

1. Cations form by losing electrons.

3. a. 1+

b. 2-

c. 2+, 3+

5. a. 1+

b. 1+, 3+

c. 1-

7. a. the potassium ion

b. the oxide ion

c. the cobalt(II) and cobalt(III) ions, respectively

9. a. the silver ion

b. the gold(I) and gold(III) ions, respectively

c. the bromide ion

11. a. magnesium chloride,  $\text{MgCl}_2$

b. iron(II) oxide,  $\text{FeO}$

- c. iron(III) oxide,  $\text{Fe}_2\text{O}_3$
13. a. copper(II) fluoride,  $\text{CuF}_2$   
b. calcium oxide,  $\text{CaO}$   
c. potassium phosphide,  $\text{K}_3\text{P}$
15. a. potassium sulfate,  $\text{K}_2\text{SO}_4$   
b. ammonium sulfide,  $(\text{NH}_4)_2\text{S}$   
c. ammonium phosphate,  $(\text{NH}_4)_3\text{PO}_4$
17. a. lead(IV) sulfate,  $\text{Pb}(\text{SO}_4)_2$   
b. sodium triiodide,  $\text{NaI}_3$   
c. lithium dichromate,  $\text{Li}_2\text{Cr}_2\text{O}_7$
19. a. silver sulfite,  $\text{Ag}_2\text{SO}_3$   
b. sodium hydrogen carbonate,  $\text{NaHCO}_3$   
c. iron(III) chlorate,  $\text{Fe}(\text{ClO}_3)_3$
21.  $\text{SO}_3$  is sulfur trioxide, while  $\text{SO}_3^{2-}$  is the sulfite ion.